

**CHEMISTRY A LEVEL PAPER 1 MARK SCHEME**

Question number	Answer	Additional guidance	Marks
1(a)	B		1
1(b)	B		1
1(c)	C		1
1(d)(i)	An explanation that makes reference to the following points: <ul style="list-style-type: none">the shape is V-shaped/bent (1)because the three electron pairs on Sn repel one another (1)	First marking point can be awarded for a correctly-drawn diagram	2
1(d)(ii)	<ul style="list-style-type: none">prediction: any value between 95 and 119° (1)justification: lone pair – bond pair repulsion is greater than bond pair – bond pair repulsion (and hence the angle is less than 120°) (1)		2

(Total for Question 1 = 7 marks)



Question number	Answer	Additional guidance	Marks
2(a)	D		1
2(b)	B		1
2(c)	D		1
2(d)	<ul style="list-style-type: none">order: GZ > DX > EY (1) Justification: <ul style="list-style-type: none">the ions in GZ have higher charges (than those in both EY and DX) (1)the ions in DX are smaller than those in EY (1)		3
2(e)	<ul style="list-style-type: none">construction of balanced cycle (1)substitution and evaluation of 2nd IE (1)	<u>Example calculation</u> $- 2258 = -590 - 2^{\text{nd}} \text{ IE} + 2 (349) - 178 - 2 (122) - 796$ <p>hence 2nd IE = (+) 1148 (kJ mol⁻¹)</p> correct answer, no working scores 2 marks	2

(Total for Question 2 = 8 marks)



Question number	Answer	Additional guidance	Marks
3(a)	<ul style="list-style-type: none"> (V) $(1s^2 2s^2 2p^6 3s^2) 3p^6 3d^3 4s^2$ (1) (V³⁺) $(1s^2 2s^2 2p^6 3s^2) 3p^6 3d^2$ (1) 	Allow $4s^2 3d^3$	2
3(b)(i)	<p>An explanation that makes reference to the following points:</p> <ul style="list-style-type: none"> E^\ominus of redox system 5 is more negative / less positive than that of both redox systems 4 and 3 (but not than that of redox systems 2 and 1) (1) therefore SO₂ releases electrons to / reduces VO₂⁺ to VO²⁺ and then VO²⁺ to V³⁺ (1) (yellow to) blue <u>then</u> green (1) 	<p>Accept explanations based on calculating E^\ominus_{cell} for the reactions</p> <p>Accept correct use of anticlockwise rule</p> <p>Ignore green colour before blue</p>	3
3(b)(ii)	<p>The reaction is not feasible because:</p> <ul style="list-style-type: none"> E^\ominus of redox system 2 is less negative than that of redox system 1 (1) therefore V²⁺ ions in system 2 will not release electrons to the V²⁺ ions in system 1 (1) 	<p>Accept explanations based on calculating E^\ominus_{cell} for the reactions</p> <p>Accept correct use of anticlockwise rule</p>	2

(Total for Question 3 = 7 marks)



Question number	Answer	Additional Guidance	Marks
4(a)	<ul style="list-style-type: none">axes: correct way round, labelled, suitable scale (1)all points plotted correctly to nearest gridline AND straight line through first 3 points, straight line through last 3 points, and the two lines intersect (1)	The scale is suitable if the distance between the first point plotted and the last point plotted covers more than half of the graph paper on each axis.	2
4(b)(i)	10 (cm ³) (1)	Award mark for a value read correctly from the candidate's graph	1
4(b)(ii)	$(0.005 \times 0.5) = 2.5(0) \times 10^{-3} / 0.0025$ (mol) (1)		1
4(b)(iii)	$(0.010 \times 1.00) = 1 \times 10^{-2} / 0.01$ (mol) (1)	Answer to (iii) csq on (ii)	1
4(b)(iv)	x = 4 and y = 2 (1)	Answer to (iv) csq on (i) and (ii), but x + y must total 6	1
4(c)	$\text{Cu}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq}) \rightarrow \text{Cu}(\text{OH})_2(\text{s})$ <ul style="list-style-type: none">correctly balanced equation (1)state symbols (1)	Allow $[\text{Cu}(\text{H}_2\text{O})_6]^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq}) \rightarrow \text{Cu}(\text{H}_2\text{O})_4(\text{OH})_2(\text{s}) + 2\text{H}_2\text{O}(\text{l})$	2

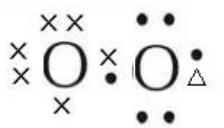
(Total for Question 4 = 8 marks)



Question number	Answer	Additional guidance	Marks
5(a)	$\text{Cl}_2 + 2\text{NaOH} \rightarrow \text{NaOCl} + \text{NaCl} + \text{H}_2\text{O}$ (1)	Accept $\text{Cl}_2 + 2\text{OH}^- \rightarrow \text{OCl}^- + \text{Cl}^- + \text{H}_2\text{O}$ Ignore state symbols	1
5(b)	An explanation that makes reference to the following points: <ul style="list-style-type: none">Cl has oxidation number +4 to in ClO_2 (1)Cl (in ClO_2) changes oxidation number to +3 (ClO_2^-) and +5 (ClO_3^-) (1)(therefore) chlorine/it (in ClO_2) has been both oxidised and reduced (1)	Accept 4+ Accept 3+ and 5+ Allow answers in any order	3
5(c)(i)	A		1
5(c)(ii)	An answer that makes reference to the following points: <ul style="list-style-type: none">Mn changes from +7 to +2 and Cl changes from -1 to 0 (1)therefore ratio is MnO_4^- to 5Cl^- /the ratio of MnO_4^- to Cl^- is 1 to 5 (1)	Accept 7+ and 2+ Accept 1-	2
5(d)(i)	An explanation that makes reference to the following points: <ul style="list-style-type: none">the HBr dissolves in water (in the air) (1)and forms droplets (of hydrobromic acid) (1)		2
5(d)(ii)	<ul style="list-style-type: none">the (conc.) H_2SO_4 is reduced to SO_2 (1)	Accept redox	1

(Total for Question 5 = 10 marks)



Question number	Answer	Additional guidance	Marks
6(a)	<ul style="list-style-type: none">determines number of mol of CO₂ produced (1)calculates number of mol NaO₂ in 880 g (1)calculates the number of mol of CO₂ reacted with NaO₂ and hence determines the number of mol of CO₂ in excess (1)calculation of mass of Na₂O = 248 g (1)	<p><u>Example of calculation</u></p> <p>2 mol glucose produces 12 mol CO₂ 880 g NaO₂ = 16 mol 16 mol NaO₂ remove 8 mol CO₂ (therefore 4 mol CO₂ is excess) mass of Na₂O required is (4 x 62) = 248 g</p> <p>3rd mark is csq on answers given in 1st and 2nd marks 4th mark is csq on answer given in 3rd mark</p> <p>Correct answer with units and no working scores 4 marks.</p>	4
6(b)	<p>An explanation that makes reference to the following points:</p> <div style="text-align: center;"></div> <ul style="list-style-type: none">it is a radical because it has an unpaired electron (1)	<p>Accept any other symbols in place of dots and crosses, including a minus sign to replace the triangle</p>	2



Question number	Answer	Additional guidance	Marks
6(c)	<ul style="list-style-type: none">• calculation of number of moles Na_2O_2 (= mols H_2O_2) (1)• calculation of number of moles Ce^{4+} (1)• determination of the ratio $\text{H}_2\text{O}_2 : \text{Ce}^{4+}$ (1)• formula of cerium ion is Ce^{3+} (1)	<p><u>Example of answer</u></p> $n(\text{Na}_2\text{O}_2) = (0.403 \div 78.0) = 5.17 \times 10^{-3}$ $n(\text{Ce}^{4+}) = (24.85 \div 1000 \times 0.420) = 1.04 \times 10^{-2}$ $n(\text{Ce}^{4+}) : n(\text{H}_2\text{O}_2) = 1.04 \times 10^{-2} : 5.17 \times 10^{-3} = 2 : 1$ <p>1st and 2nd mark: accept answers to any significant figures, except 1s.f</p> <p>3rd mark csq on answers given in 1st and 2nd marks</p> <p>4th mark csq on answer given in 3rd mark</p>	4

(Total for Question 6 = 10 marks)



Question number	Answer	Additional guidance	Marks
7(a)	B		1
7(b)(i)	<ul style="list-style-type: none"> • calculation of $n(\text{CO})$ at equilibrium and $n(\text{H}_2)$ at equilibrium (1) • converting number of moles to concentration (1) • evaluation of K_c by substitution (1) • correct answer to 3 sf (1) • units: $\text{dm}^6 \text{ mol}^{-2}$ (1) 	<p><u>Example of calculation</u> $n(\text{CO})$ at equilibrium = 0.114 (mol) $n(\text{H}_2)$ at equilibrium = 0.228 (mol) $[\text{CO}] = 0.0950 \text{ mol dm}^{-3}$, $[\text{H}_2] = 0.190 \text{ mol dm}^{-3}$, $[\text{CH}_3\text{OH}] = 0.0717 \text{ mol dm}^{-3}$ $K_c = 0.0717 \div (0.0950 \times 0.190^2) = 20.9068\dots$ $= 20.9 \text{ dm}^6 \text{ mol}^{-2}$</p> <p>Allow $\text{mol}^{-2} \text{ dm}^6$</p> <p>3rd and 4th marks csq on answers given in 1st and 2nd marks</p> <p>Correct final answer to 3 sf with units but no working scores 5 marks</p>	5
7(b)(ii)	<p>An explanation that makes reference to the following points:</p> <ul style="list-style-type: none"> • an increase in temperature shifts the equilibrium to the left (1) • an increase in pressure shifts the equilibrium to the right (1) • these changes produce opposing effects, so to predict the effect on the yield it is necessary to know the relative effects of each one (1) 		3

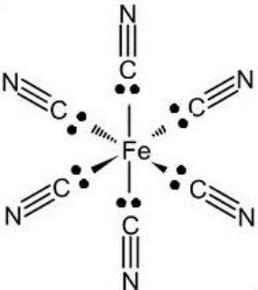
(Total for Question 7 = 9 marks)



Question number	Answer	Additional guidance	Marks
8(a)	A		1
8(b)(i)	<p>An explanation that makes reference to the following points:</p> <ul style="list-style-type: none"> stage 2 is an equilibrium reaction / only partial ionisation occurs (1) therefore fewer hydrogen ions are formed (1) 	Accept dissociation for ionisation	2
8(b)(ii)	<ul style="list-style-type: none"> rearrangement of equation $\text{pH} = -\log [\text{H}^+]$ and substitution to give final answer (1) 	<p>Example calculation: $[\text{H}^+] = 10^{-\text{pH}}$ $10^{-0.97} = 0.107 \text{ (mol dm}^{-3}\text{)}$</p> <p>Allow 0.11 (mol dm⁻³)</p> <p>Correct answer with no working scores 1 mark</p>	1
8(c)	<ul style="list-style-type: none"> rearrangement of K_a expression (1) calculation of $[\text{CH}_3\text{COOH}]$ and $[\text{CH}_3\text{COO}^-]$ (1) substitution, and evaluation of $[\text{H}^+]$ in the buffer solution (1) conversion of $[\text{H}^+]$ to pH for buffer solution (1) 	<p>Example of calculation :</p> $[\text{H}^+] = K_a \frac{[\text{CH}_3\text{COOH}]}{[\text{CH}_3\text{COO}^-]}$ <p>$[\text{CH}_3\text{COOH}] = 0.333 \text{ mol dm}^{-3}$ and $[\text{CH}_3\text{COO}^-] = 0.167 \text{ mol dm}^{-3}$</p> <p>$[\text{H}^+] = 1.74 \times 10^{-5} \times 0.333 / 0.167$ $= 3.48 \times 10^{-5} \text{ (mol dm}^{-3}\text{)}$ so $\text{pH} = -\lg 3.48 \times 10^{-5} = 4.46$</p> <p>Accept answers that use forms of the Henderson-Hasselbach equation</p> <p>Correct answer with no working scores 4 marks</p>	4

(Total for Question 8 = 8 marks)



Question number	Answer	Additional guidance	Marks
9(a)(i)	A		1
9(a)(ii)	 <ul style="list-style-type: none">• 3D shape correctly shown (1)• electron pairs shown OR all six bonds shown clearly to the carbons (1)	Charge and square brackets not required Allow dotted lines for the wedges going backwards Allow arrows to represent dative covalent bonds instead of electron pairs	2
9(b)	<ul style="list-style-type: none">• step 1 – ligand substitution/ligand exchange (1)• step 2 – redox (1)	Allow reduction and oxidation	2
9(c)	<ul style="list-style-type: none">• identification: green precipitate is iron(II) hydroxide (1) <p>Explanation:</p> <ul style="list-style-type: none">• precipitate turns brown because iron(II) hydroxide is oxidised by oxygen (and water) in the air (1)• to form iron(III) hydroxide (1)	Allow $\text{Fe}(\text{OH})_2$ / $\text{Fe}(\text{H}_2\text{O})_4(\text{OH})_2$ Allow $\text{Fe}(\text{OH})_3$ / $\text{Fe}(\text{H}_2\text{O})_3(\text{OH})_3$	3

(Total for Question 9 = 8 marks)



Question number	Answer	Additional guidance	Marks
10(a)	An explanation that makes reference to the following points: <ul style="list-style-type: none"> they get more stable down the group (1) because the size of the cations increases/charge density of cations decreases (1) and so carbonate ions are less polarised (1) 		3
10(b)	<ul style="list-style-type: none"> rearrangement of equation (1) calculation of $\Delta H_{\text{solution}}$ (1) 	<u>Example of calculation</u> $-2493 + \Delta H_{\text{solution}} = -1920 + (-2 \times 364)$ $\Delta H_{\text{solution}} = -155 \text{ (kJ mol}^{-1}\text{)}$ Correct sign must be given in final answer Correct answer and sign with no working scores 2 marks	2
10(c)(i)	An explanation that makes reference to the following points: <ul style="list-style-type: none"> breaking the lattice is endothermic and the hydration of ions is exothermic (1) (therefore the dissolving of magnesium sulphate is exothermic) because the enthalpy of hydration (of the ions) is greater in magnitude than the lattice energy (of MgSO_4) (1) 		2
10(c)(ii)	<ul style="list-style-type: none"> $\Delta G^\ominus = -87 - (298 \times -0.210)$ $= -24(.42) \text{ (kJ mol}^{-1}\text{)}$ (1) since ΔG is negative the process/reaction is spontaneous/feasible (1) 		2



Question number	Answer	Additional guidance	Marks												
*10(d)	<p>This question assesses a student's ability to show a coherent and logically structured answer with linkages and fully-sustained reasoning.</p> <p>Marks are awarded for indicative content and for how the answer is structured and shows lines of reasoning.</p> <p>The following table shows how the marks should be awarded for indicative content.</p> <table border="1" data-bbox="398 584 853 984"><thead><tr><th data-bbox="398 584 607 799">Number of indicative marking points seen in answer</th><th data-bbox="607 584 853 799">Number of marks awarded for indicative marking points</th></tr></thead><tbody><tr><td data-bbox="398 799 607 839">6</td><td data-bbox="607 799 853 839">4</td></tr><tr><td data-bbox="398 839 607 879">5-4</td><td data-bbox="607 839 853 879">3</td></tr><tr><td data-bbox="398 879 607 919">3-2</td><td data-bbox="607 879 853 919">2</td></tr><tr><td data-bbox="398 919 607 959">1</td><td data-bbox="607 919 853 959">1</td></tr><tr><td data-bbox="398 959 607 984">0</td><td data-bbox="607 959 853 984">0</td></tr></tbody></table>	Number of indicative marking points seen in answer	Number of marks awarded for indicative marking points	6	4	5-4	3	3-2	2	1	1	0	0	<p>Guidance on how the mark scheme should be applied:</p> <p>The mark for indicative content should be added to the mark for lines of reasoning. For example, an answer with five indicative marking points that is partially structured with some linkages and lines of reasoning scores 4 marks (3 marks for indicative content and 1 mark for partial structure and some linkages and lines of reasoning).</p> <p>If there are no linkages between points, the same five indicative marking points would yield an overall score of 3 marks (3 marks for indicative content and no marks for linkages).</p>	6
Number of indicative marking points seen in answer	Number of marks awarded for indicative marking points														
6	4														
5-4	3														
3-2	2														
1	1														
0	0														



Question number	Answer	Additional guidance	Marks								
*10(d) cont.	<p>The following table shows how the marks should be</p> <table border="1" data-bbox="398 236 1265 746"> <thead> <tr> <th data-bbox="398 236 943 419"></th> <th data-bbox="943 236 1265 419">Number of marks awarded for structure of answer and sustained line of reasoning</th> </tr> </thead> <tbody> <tr> <td data-bbox="398 419 943 563">Answer shows a coherent and logical structure with linkages and fully sustained lines of reasoning demonstrated throughout.</td> <td data-bbox="943 419 1265 563">2</td> </tr> <tr> <td data-bbox="398 563 943 675">Answer is partially structured with some linkages and lines of reasoning.</td> <td data-bbox="943 563 1265 675">1</td> </tr> <tr> <td data-bbox="398 675 943 746">Answer has no linkages between points and is unstructured.</td> <td data-bbox="943 675 1265 746">0</td> </tr> </tbody> </table> <p>awarded for structure and lines of reasoning.</p> <p>Indicative content</p> <p>$(\Delta G^{\circ}_{\text{solution}} = \Delta H^{\circ}_{\text{solution}} - T\Delta S^{\circ}_{\text{system}})$</p> <ul style="list-style-type: none"> • for BaSO₄: $\Delta H^{\circ}_{\text{solution}}$ and $-T\Delta S^{\circ}_{\text{system}}$ are both positive (1) • for CaSO₄: $\Delta H^{\circ}_{\text{solution}}$ is negative and $-T\Delta S^{\circ}_{\text{system}}$ is positive (1) • but the magnitude of $-T\Delta S^{\circ}_{\text{system}}$ is greater than that of $\Delta H^{\circ}_{\text{solution}}$ (1) • therefore $\Delta G^{\circ}_{\text{solution}}$ for both salts is positive (1) • when $\Delta G^{\circ}_{\text{solution}}$ is positive the salt is only slightly soluble (1) • BaSO₄ is less soluble than CaSO₄ because $\Delta G^{\circ}_{\text{solution}}$ is more positive (1) 		Number of marks awarded for structure of answer and sustained line of reasoning	Answer shows a coherent and logical structure with linkages and fully sustained lines of reasoning demonstrated throughout.	2	Answer is partially structured with some linkages and lines of reasoning.	1	Answer has no linkages between points and is unstructured.	0		
	Number of marks awarded for structure of answer and sustained line of reasoning										
Answer shows a coherent and logical structure with linkages and fully sustained lines of reasoning demonstrated throughout.	2										
Answer is partially structured with some linkages and lines of reasoning.	1										
Answer has no linkages between points and is unstructured.	0										

(Total for Question 10 = 15 marks)

