



Definitions and Concepts for Edexcel Chemistry A-level

Topic 10: Equilibrium 1

Features of a dynamic equilibrium:

- Reversible reaction in a closed system.
- Forward and backward reactions occur simultaneously.
- Rate of forward reaction = Rate of backward reaction.
- There is **no net change** in concentrations of reactants and products (see above).

Homogeneous system: A system, where all the chemicals are in the same phase.

Heterogeneous system: A system, where not all the chemicals are in the same phase.

Le Chatelier's Principle: If there's a small change in concentration, pressure/volume or temperature (disruption to equilibrium) to a closed system in dynamic equilibrium, the equilibrium will shift to minimise the change (to restore the equilibrium).

Contact process: A method of production of H₂SO₄. The key step involves the conversion of SO₂ to SO₃ in presence of oxygen, catalysed by V₂O₅. An example of heterogeneous catalysis.

